

# 12 1 Stoichiometry Study Guide

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Objectives 12.1.1 Explain how balanced equations apply to both chemistry and everyday situations.

12.1.2 Interpret balanced chemical equations in terms of moles, representative particles, mass, and gas volume at STP. 12.1.3 Identify the quantities that are

always conserved in chemical reactions. Guide for Reading 12.1 The Arithmetic of Equations 12 Related

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Stoichiometry Practice Questions Use the link below to answer the following questions: 1. What does stoichiometry help you figure out? 2. What are all reactions dependent upon? 3. If I have ten hydrogen molecules and three oxygen molecules, how many molecules of water can I make? Chem Int CC Ch 12 - Stoichiometry - Answers (09.15).pdf ... TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between ... VIBRATIONS AND WAVES CHAPTER 12

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